

ABSORPTION SPECTRUM OF Ni(II) IONS DOPED IN LITHIUM SODIUM POTASSIUM SULPHATE SINGLE CRYSTAL

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The optical absorption spectra of lithium sodium potassium sulphate doped Ni^{2+} ions at 300 and 77 K was recorded. The observed bands are assigned as transitions from the ground state ${}^3A_{2g}(F)$ to various excited triplet and singlet states of the Ni^{2+} ion in octahedral symmetry. The splittings observed for the bands ${}^3T_{1g}(F)$ and ${}^3T_{2g}(F)$ at liquid nitrogen temperature have been explained as due to spin-orbit interaction. The oscillator strengths of the transitions were computed from the area under the curves. From a detailed analysis of the bands of the spectrum, the crystal field parameters B , C , Dq and ξ are evaluated.

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1. Introduction

The optical absorption spectra of Ni^{2+} ion in various environments have been studied extensively by several authors [1-6]. Since Ni^{2+} ion among the transition metal ions has strong absorption bands in the 200-1500 nm region, which are well assigned, Ni^{2+} ion was used as an impurity and optical probe to study the electronic absorption spectra in different host lattices with different symmetry ligands [7-11]. However, there seem to be no reports available in the literature on optical studies of Ni^{2+} ions in the lithium sodium potassium sulphate (LSPS) lattice. Hence the studies of Ni^{2+} ion in LSPS were undertaken at 300 and 77 K to obtain the crystal field parameters and to estimate the site symmetry of the ion in the host lattice.

The triple salt, lithium sodium potassium sulphate ($\text{Li}_2\text{NaK}(\text{SO}_4)_2$) has been the subject of many investigations for the past few years [12-18]. This compound is known to crystallize in an orthorhombic system with space group $P2_12_12_1$ and lattice dimensions, $a = 4.95$, $b = 7.81$ and $c = 19.02$ Å [16].

2. Experimental

Single crystals of LSPS were grown by slow evaporation of saturated aqueous solution of the corresponding sulphates mixed in their equimolar ratio. The doping of Ni^{2+} ions was made by adding about 0.1 mole percent of nickel sulphate to the saturated solution. The crystals grown were clear and green in colour.

The optical absorption spectrum was recorded on a Cary-17D spectrophotometer both at room (300 K) and liquid nitrogen (77 K) temperatures. The oscillator strengths were computed by measuring the area under the curves of the spectra.

3. Theory

Ni^{2+} has eight d electrons. The electronic states of Ni^{2+} in a strong octahedral field arising from the ground and the excited electronic configurations are as follows:

$$\begin{array}{ll} \text{from the ground} & (t_{2g})^6 (e_g)^2 : {}^3A_{2g}, {}^1A_{1g}, {}^1E_g, \\ \text{from the first excited} & (t_{2g})^5 (e_g)^3 : {}^3T_{1g}, {}^3T_{2g}, {}^1T_{1g}, {}^1T_{2g}, \\ \text{from the second excited} & (t_{2g})^4 (e_g)^4 : {}^3T_{1g}, {}^1A_{1g}, {}^1E_g, {}^1T_{2g}. \end{array}$$

According to Hund's rule, ${}^3A_{2g}$ lies lowest and forms the ground state of the ion. The energy expressions for all the states in the absence of spin-orbit interaction were given in the form of matrices by Tanabe and Sugano [19]. Lakshman et al. [20] presented these matrices in a linear form. In the presence of spin-orbit interaction, the energy levels are designated as $\Gamma_1, \Gamma_2, \Gamma_3, \Gamma_4$ and Γ_5 . The energy matrices for these spin-orbit levels were presented by Liehr and Ballhausen [21] in terms of F_2, F_4, Dq and λ and in terms of B, C, Dq and ξ by Lakshman and Rao [22].

4. Results and analysis

The crystal shows the characteristic absorption of Ni^{2+} ions in the visible and near infrared regions. The optical absorption spectrum observed at room temperature is shown in Fig. 1. Five bands have been observed at room temperature, one in the near infrared at 8850 cm^{-1} , three in the visible at 13889, 14793 (and 15060), 22222 cm^{-1} and one in the near ultraviolet at 26596 cm^{-1} . Among these five bands, the band in the near ultraviolet is the most intense.

On cooling the crystal to liquid nitrogen temperature, changes in intensity and band positions are observed. The spectrum observed at liquid nitrogen temperature is shown in Fig. 2. From the spectra, it can be seen that at low temperature the band 8850 cm^{-1} at 300 K, has resolved into two bands with band maxima at 9174 and 9524 cm^{-1} , with a blue shift of 499 cm^{-1} . The band at 14927 cm^{-1} has been found to split into four components with maxima at 15015, 15152, 15748 and 16026 cm^{-1} , while the bands 13889, 22222 and 26596 cm^{-1} are shifted to 13986, 22831 and 27174 cm^{-1} , respectively.

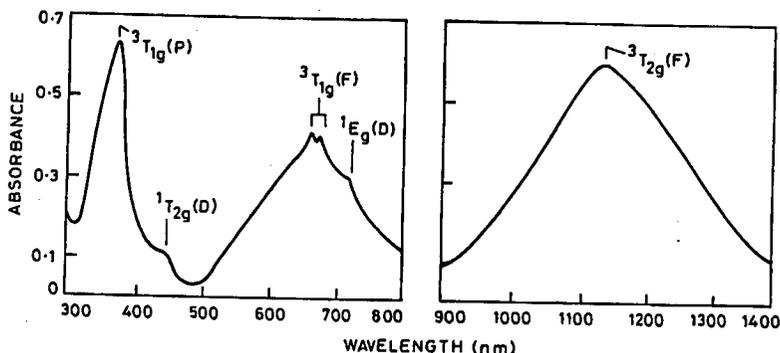


Fig. 1. Absorption spectrum of Ni²⁺ in lithium sodium potassium sulphate at room temperature (300 K).

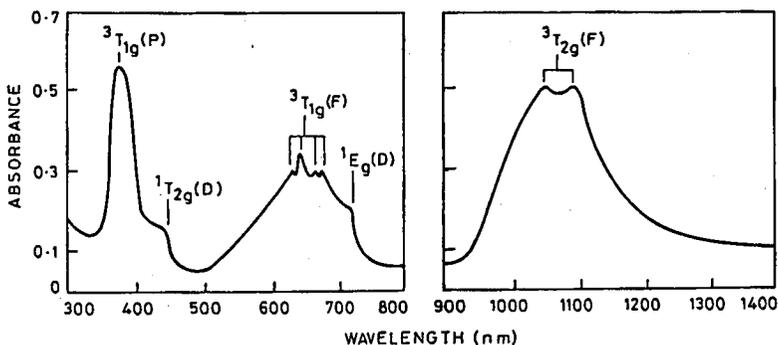


Fig. 2. Absorption spectrum of Ni²⁺ in lithium sodium potassium sulphate at liquid nitrogen temperature (77 K).

From the nature and observed position of the bands, they have been attributed to an ion of Ni²⁺ in octahedral symmetry. The ground state electronic configuration of the Ni²⁺ ion in octahedral symmetry is ${}^3A_{2g}(F)$. According to the energy level scheme, three spin-allowed bands should be observed which arise due to the transitions from the ground state to the excited states ${}^3T_{2g}(F)$, ${}^3T_{1g}(F)$, and ${}^3T_{1g}(P)$ arranged in the order of increasing energy.

Often most of the nickel complexes show a double-peaked absorption at room temperature [3, 23]. Indeed, in the present work also, we observed a double-peaked band separated by 267 cm^{-1} located at 14793 and 15060 cm^{-1} . The intense bands observed at 8850 and 26596 cm^{-1} have therefore been attributed to ${}^3A_{2g}(F) \rightarrow {}^3T_{2g}(F)$ and ${}^3A_{2g}(F) \rightarrow {}^3T_{1g}(P)$ transitions, respectively, and the double-peaked band is attributed to ${}^3A_{2g}(F) \rightarrow {}^3T_{1g}(F)$ transition. The three spin-allowed bands are expected to show blue shift at low temperature, as their corresponding states ${}^3T_{2g}(F)$, ${}^3T_{1g}(F)$ and ${}^3T_{1g}(P)$ have positive slopes in the Tanabe-Sugano energy level diagram [19], given for d^8 configuration. Therefore, the observed blue shifts for the ${}^3T_{2g}(F)$ and ${}^3T_{1g}(P)$ bands at low temperature are in accordance with

the theory. In the present work, no such observation could be made for the ${}^3T_{1g}(F)$ band as it has been found to split on cooling the crystal to liquid nitrogen temperature. Such splittings in ${}^3T_{1g}(F)$ band have been reported in literature by several investigators [5, 6, 24, 25]. This characteristic splitting of the band further supports the assignment. The double-peaked band attributed to ${}^3A_{2g}(F) \rightarrow {}^3T_{1g}(F)$ splits into four components and the near infrared band attributed to ${}^3A_{2g}(F) \rightarrow {}^3T_{2g}(F)$ into two. Due to spin-orbit interaction, each of these bands must split into four components. But in the present investigation, the near infrared band is not properly resolved and only two components are observed. Similar features are reported for Ni^{2+} ions in various host lattices [24, 26].

Lever [23] noted that ${}^3T_{1g}(F)$ appears as well defined double-peaked band if Dq/B ratio is near unity. The calculated Dq/B ratio at room temperature in the present work is also nearer to unity (0.91). This also justifies the assignment of the double-peaked band to the transition from ${}^3A_{2g}(F) \rightarrow {}^3T_{1g}(F)$.

According to the theory, some spin-forbidden triplet-to-singlet transitions may also appear and the intensity of such bands would be lower. The bands observed at 13889 and 22222 cm^{-1} are assigned to ${}^3A_{2g}(F) \rightarrow {}^1E_g(D)$ and ${}^3A_{2g}(F) \rightarrow {}^1T_{2g}(D)$ transitions, respectively. On cooling the crystal to liquid nitrogen temperature, the above two spin-forbidden bands show blue shifts as their energy levels have positive slopes in the Tanabe-Sugano energy level diagram.

5. Discussion

In optical studies, one important and interesting feature observed on cooling to low temperature is the splitting of some of the bands. The reason for this splitting may be spin-orbit interaction or lowering of symmetry or superposition of certain vibrational modes of radicals or ligands. If the bands are split due to lowering of symmetry, the other orbital doublet and triplet states are also expected to split into various components. This is not observed in the present work. The intensity of the splitting of the bands observed due to simultaneous electronic and vibrational transitions is very weak [1]. In the present work, the average intensity of the split bands is not weak. Therefore the nature of the splittings observed at liquid nitrogen temperature for the ${}^3T_{1g}(F)$ and ${}^3T_{2g}(F)$ bands appear to be due to the spin-orbit effect. Therefore, the ligand field calculations are carried out in terms of the octahedral field combined with spin-orbit coupling. The ${}^3T_{1g}(F)$ level splits into three ($\Gamma_3 + \Gamma_5$, Γ_4 and Γ_1) and four (Γ_5 , Γ_3 , Γ_4 and Γ_1) component levels, respectively, in the first and second order configurational interactions. Using the free ion spin-orbit coupling parameter $\xi = 600 \text{ cm}^{-1}$, the calculated first order separation between the components of ${}^3T_{1g}(F)$ state would be 140 and 280 cm^{-1} . Since the observed splittings of 137 cm^{-1} (15152–15015) and 278 cm^{-1} (16026–15748) are approximately of the same order of spin-orbit splitting expected for Ni^{2+} ion in the crystal, the energy matrices inclusive of spin-orbit effect are diagonalized for different values of Dq , with $B = 980 \text{ cm}^{-1}$, $C = 3280 \text{ cm}^{-1}$ and $\xi = 600 \text{ cm}^{-1}$. The best fit of the observed bands at 77 K could be obtained with $Dq = 930 \text{ cm}^{-1}$. The corresponding energy level diagram (E vs. Dq) is shown in Fig. 3. The observed and calculated band maxima positions along with their oscillator strengths are presented in Table I.

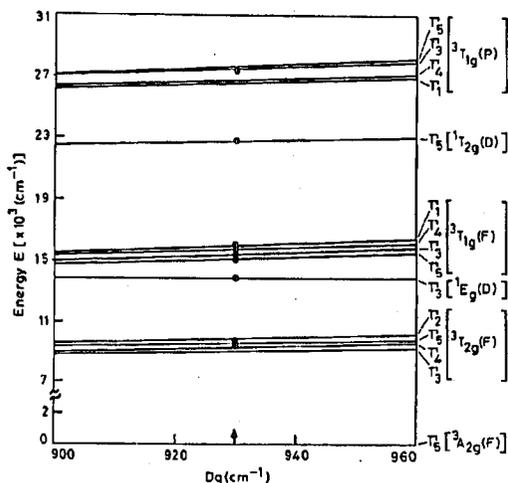


Fig. 3. Energy level diagram of Ni^{2+} in lithium sodium potassium sulphate in cubic environment plotted as a function of crystal field parameter Dq , with $B = 980 \text{ cm}^{-1}$, $C = 3280 \text{ cm}^{-1}$ and $\xi = 600 \text{ cm}^{-1}$. The solid circles show the experimental values at 77 K.

TABLE I

The observed and calculated energies, oscillator strengths, and assignments for the bands of Ni^{2+} in lithium sodium potassium sulphate ($B = 980 \text{ cm}^{-1}$, $C = 3280 \text{ cm}^{-1}$, $Dq = 930 \text{ cm}^{-1}$ and $\xi = 600 \text{ cm}^{-1}$).

Transition from	Spin-orbit designation	Band positions [cm^{-1}]			Oscillator strengths* $f (\times 10^{-5})$	
		Cal.	Obs.		300 K	77 K
			300 K	77 K		
${}^3A_{2g}(F)$	Γ_5					
${}^3T_{2g}(F)$	Γ_3	9067	8850			
	Γ_4	9210		9174	5.71	3.76
	Γ_5	9572		9524		
	Γ_2	9690				
${}^1E_g(D)$	Γ_3	13879	13889	13986	—	—
${}^3T_{1g}(F)$	Γ_5	15197	14793	15015		
	Γ_3	15411	15060	15152	7.31	5.41
	Γ_4	15739		15748		
	Γ_1	15907		16026		
${}^1T_{2g}(D)$	Γ_5	22812	22222	22831	—	—
${}^3T_{1g}(P)$	Γ_1	26658	26596			
	Γ_4	26720				
	Γ_3	27593		27174	12.2	7.22
	Γ_5	27693				

* Sum of all split components.

When the crystal is cooled from room temperature to liquid nitrogen temperature, all the bands showed a decrease in intensity, which is characteristic of the vibronic intensity mechanism and is expected for the $d-d$ transitions of an octahedral transition metal ion complex. Similar observations were reported by Venkata Subbaiah et al. [6], Lakshman and Rao [22], McPherson and Devaney [27] and Rao et al. [28] in their absorption studies of Ni^{2+} complexes.

The energy parameters (B , C and Dq) obtained in the present work are given in Table II along with the parameters reported for Ni^{2+} ion in various crystals. From the table it can be seen that the energy parameters evaluated in the present work are in good agreement with the values reported by several investigators for Ni^{2+} ion in O_h symmetry associated with the spin-orbit interaction. This also strengthens our assignments and analysis on the basis of spin-orbit interaction.

The interelectronic repulsion parameter B for free Ni^{2+} ion is 1080 cm^{-1} [33]. In the present work, we obtained a B value equal to 980 cm^{-1} at 77 K and this suggests that the ionic bonding is more predominant in the complex.

TABLE II
Comparison of energy parameters of Ni^{2+} in LSPS with those reported for other lattices.

Ni^{2+} ion in	B [cm^{-1}]	C [cm^{-1}]	Dq [cm^{-1}]	Ref.
$MgTi_2(SO_4)_2 \cdot 6H_2O$	900	3660	960	6
$Cd(NH_4)_2(SO_4)_2 \cdot 6H_2O$	955	4202	922	24
$MgKPO_4 \cdot 6H_2O$	890	3740	935	25
$MgNH_4PO_4 \cdot 6H_2O$	910	3981	915	26
$ZnTi_2(SO_4)_2 \cdot 6H_2O$	905	3420	950	28
$LiCsSO_4$	890	3600	950	29
$NaNH_4SO_4 \cdot 2H_2O$	740	3160	1000	30
$LiGa_5O_8$	928	3073	977	31
$K_2Zn(SO_4)_2 \cdot 6H_2O$	944	3411	941	31
$CdK_2(SO_4)_2 \cdot 6H_2O$	940	3760	937	32
$Li_2NaK(SO_4)_2$	980	3280	930	present work

6. Conclusions

The optical absorption spectra of LSPS: Ni^{2+} lattice has been studied at 300 and 77 K. From the result and discussion, the transitions of Ni^{2+} in LSPS lattice were assigned and from the analysis of the spectra, the crystal field (Dq), Racah (B and C) and spin-orbit interaction (ξ) parameters were evaluated.

Basing on the results and analysis of this investigation, it is reasonable to assume that the spectrum is characteristic of Ni^{2+} ion in octahedral site symmetry associated with spin-orbit interaction.

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References

- [1] T.S. Piper, N. Koertge, *J. Chem. Phys.* **32**, 559 (1960).
- [2] J. Ferguson, H.J. Guggenheim, H. Kamimura, Y. Tanabe, *J. Chem. Phys.* **42**, 775 (1965).
- [3] E.I. Solomon, C.J. Ballhausen, *Mol. Phys.* **29**, 279 (1975).
- [4] D.R. Rosseinsky, I.A. Dorrity, *Coord. Chem. Rev.* **25**, 31 (1978).
- [5] B. Ghosh, R.K. Mukherjee, *Phys. Status Solidi B* **102**, 189 (1980).
- [6] A. Venkata Subbaiah, J. Lakshmana Rao, R. Murali Krishna, S.V.J. Lakshman, *Z. Nat.forsch. A* **47**, 813 (1992).
- [7] W. Low, *Phys. Rev.* **109**, 247 (1958).
- [8] H. Hartmann, H. Muller, *Disc. Farad. Soc.* **26**, 49 (1958).
- [9] O.G. Holmes, D.S. McClure, *J. Chem. Phys.* **26**, 1686 (1957).
- [10] A.S. Chakravarty, R. Chatterjee, *Ind. J. Phys.* **35**, 643 (1961).
- [11] D.S. McClure, *J. Phys. Chem. Solids* **3**, 311 (1957).
- [12] I.N. Lepeshkov, N.V. Bodaleva, L.T. Kotava, *Zh. Neorg. Khim.* **3**, 2781 (1958).
- [13] I.N. Lepeshkov, N.V. Bodaleva, L.T. Kotava, *Russ. J. Inorg. Chem.* **6**, 864 (1961).
- [14] K. Kitahama, R. Frech, *J. Chem. Phys.* **82**, 720 (1985).
- [15] K. Kitahama, R. Frech, *Solid State Ion.* **21**, 63 (1986).
- [16] M.J. Heeg, A. Hurd, *Acta Crystallogr. C* **43**, 161 (1987).
- [17] J.K. Liang, X.T. Xu, Z. Chai, *J. Solid State Chem.* **76**, 270 (1988).
- [18] R. Rama Kumar, B.C. Venkata Reddy, *Cryst. Res. Technol.* **28**, 535 (1993).
- [19] Y. Tanabe, S. Sugano, *J. Phys. Soc. Japan* **9**, 753 (1954).
- [20] S.V.J. Lakshman, J. Lakshmana Rao, *Ind. J. Pure Appl. Phys.* **10**, 497 (1972).
- [21] A.D. Liehr, C.J. Ballhausen, *Ann. Phys. (USA)* **6**, 134 (1959).
- [22] S.V.J. Lakshman, J. Lakshmana Rao, *Spectrochim. Acta A* **35**, 703 (1979).
- [23] A.B.P. Lever, *Inorganic Electronic Spectroscopy*, 2nd ed., Elsevier, Amsterdam 1984, p. 507.
- [24] P. Sathyasainath Prasad, S. Radhakrishna, *Solid State Commun.* **64**, 561 (1987).
- [25] P. Sivaprasad, K. Ramesh, Y.P. Reddy, *Solid State Commun.* **73**, 239 (1990).
- [26] B.V.R. Chowdari, Y. Ravisekhar, C. Muralikrishna, *Phys. Status Solidi B* **109**, 315 (1982).
- [27] G.L. McPherson, K.O. Devaney, *Inorg. Chem.* **16**, 1565 (1977).
- [28] J.L. Rao, M.R. Reddy, S.V.J. Lakshman, *Phys. Status Solidi B* **151**, 599 (1989).
- [29] B.C. Venkata Reddy, B. Munibhadraiah, *Ind. J. Pure Appl. Phys.* **24**, 577 (1986).
- [30] S.V.J. Lakshman, T.V. Krishna Rao, *Pramana* **20**, 137 (1983).
- [31] D.T. Sviridov, S.D. Sviridov, *Zh. Prikl. Spektrosk.* **49**, 851 (1988).
- [32] P. Sathyasainath Prasad, S. Radhakrishna, *Cryst. Latt. Def. Amorph. Mater.* **14**, 85 (1987).
- [33] B.N. Figgis, *Introduction to Ligand Fields*, Interscience, New York 1966, p. 52.