

Magnetic Properties of Mn-Doped Amorphous SiO₂ Matrix

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Samples of Mn-doped amorphous SiO₂ matrix with manganese concentration 0.7 and 3 at.% have been prepared by a sol-gel method. Transmission electron microscopy analysis has shown that the samples contain agglomerates of amorphous silica particles 10 – 20 nm in size. Two types of Mn-rich particles are dispersed in silica matrix, smaller nanoparticles with dimensions between 3 and 10 nm, and larger crystalline areas consisting of aggregates of the smaller nanoparticles. High-temperature magnetic susceptibility reveals that dominant magnetic phase at higher temperatures is λ -MnO₂. At temperatures below $T_C = 43$ K strong ferrimagnetism originating from the minor Mn₃O₄ phase masks the relatively weak magnetism of λ -MnO₂. Magnetic field dependence of the maximum in the zero-field-cooled magnetization for both the samples in the vicinity of 40 K, and a frequency shift of the real component of the AC magnetic susceptibility in the sample with 3 at.% Mn suggest that the magnetic moments of the smaller Mn₃O₄ nanoparticles with dimensions below 10 nm are subject to thermally activated blocking process just below the Curie temperature T_C . The low-temperature maximum in the zero-field-cooled magnetization observed for both the samples below 10 K indicates possible spin glass freezing of the magnetic moments in the geometrically frustrated Mn sublattice of the λ -MnO₂ crystal structure.

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1. Introduction

Manganese oxide materials have received much attention because of their importance in technological applications such as catalysts, sensors, new magnetic materials [1]. Among them, manganese dioxide has attracted considerable interest because of possibility of numerous applications in catalysis, ion exchange, magnetoresistance devices and particularly energy storage devices [2–4]. Several polymorphs of MnO₂ has been identified so far, α , β , γ and λ types. Low dimensional Mn-oxide structures greatly affect their physical and chemical properties and improve potential of these materials for applications in various fields. In this work two manganese oxides in the Mn-rich particles embedded in the amorphous SiO₂ matrix were identified, major phase λ -MnO₂ and minor phase Mn₃O₄. λ -MnO₂ is a metastable form of manganese dioxide whose thermodynamically stable structure is β -MnO₂ [5]. λ -MnO₂ phase is usually obtained by removal of Li from LiMn₂O₄ by acid leaching or by electrochemical procedure [5, 6]. At higher temperatures Li_xMn₂O₄ spinel has paramagnetic properties. At lower temperatures a complex antiferromagnetic (AFM) ordering in λ -MnO₂ [5], and very complex magnetic ordering in LiMn₂O₄ [7] were reported. In addition, spin glass behavior was also observed at low-temperatures in λ -MnO₂, LiMn₂O₄ and Li₂Mn₂O₄ [6]. Li_xMn₂O₄ mate-

rial has attracted a wide attention because of its utilization as cathode material for rechargeable Li batteries [3–4]. In the samples studied in the present work Mn₃O₄ structure as a minor phase was found to exist simultaneously with the λ -MnO₂ phase. The synthesis procedure enabled us to obtain Mn₃O₄ nanocrystallites with dimensions below 10 nm and to study their magnetic properties characteristic to very small nanoparticles.

2. Experimental details

The silica sol with very strictly determined size and uniformity of silica sol particles was synthesized using a water glass solution of Na₂O·3SiO₂ with the silica module SiO₂/Na₂O ratio of 3.75. Then, the second solution with previously stoichiometrically adjusted quantities of Mn(NO₃)₂ · 6H₂O salt was added to silica sol solution equivalent to content of 1 and 5 at.% Mn (nominal manganese concentrations).

The content of manganese in the two samples determined by atomic absorption method is 6.36 mg and 27.34 mg per gram of the sample which corresponds to 0.7 and 3 at.% of Mn respectively. Powder X-ray diffraction (XRD) spectra were recorded on a Philips PW 1050 diffractometer using Cu K α radiation. For transmission electron microscopy (TEM) measurements the samples were deposited on a copper-grid-supported perforated transparent carbon foil. High-resolution transmission electron microscopy (HRTEM) and (semi)quantitative energy dispersive X-ray spectroscopy (EDXS) were car-

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ried out using a field-emission electron-source transmission electron microscope JEOL 2010 F operated at 200 kV. The microscope was equipped with an energy dispersive X-ray spectrometer (LINK ISIS EDS 300). The magnetization measurements were performed on a SQUID magnetometer (MPMS XL-5, Quantum Design) equipped with an AC option. The DC magnetization, measured at applied magnetic fields from 50 to 10000 Oe under zero-field-cooled (ZFC) and field-cooled (FC) regime, was recorded in the temperature range from 2 to 300 K. The AC magnetization measurements were made under an AC exciting field with amplitude of 6.5 Oe using different frequencies in the range 1-1488 Hz.

3. Results and discussion

XRD spectrum of the sample with 3 at.% Mn is presented in Fig. 1. The XRD pattern is characterized by a broad band with a maximum at $2\theta = 21.9^\circ$. The shape of the diffractogram and position of the wide band correspond to amorphous silica [8]. No separated narrow XRD lines were observed.

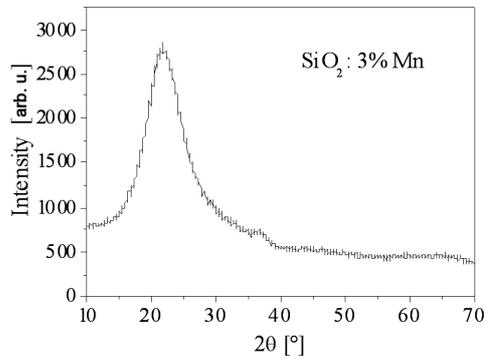


Fig. 1. X-ray diffraction pattern of the SiO₂ : 3 at.% Mn sample.

TEM image of the sample with 3 at.% Mn is shown in Fig. 2a. This image reveals that the sample contains agglomerates of globular amorphous silica particles of approximately 10-20 nm in size. EDXS spectra taken from the majority of the individual silica particles showed only Si and O suggesting that Mn is segregated inside the silica matrix. Inspection of the sample using HRTEM showed two types of Mn-rich nanoparticles dispersed inside the silica matrix:

(i) Smaller Mn-rich nanoparticles with sizes between 3 and 10 nm (marked with circles in Fig. 2b). The nanoparticles are crystalline, as can be seen from regular periodicity of lattice fringes in the HRTEM images (Fig. 2c).

(ii) Rarely, larger crystalline areas, most probably consisting of aggregates of the smaller Mn-rich nanoparticles, were also observed, Fig. 2d.

Some interplanar lattice spacings have been measured for the dispersed nanoparticles and those inside the larger crystalline areas: 2.1 Å, 2.3 Å, 2.47 Å, 2.7 Å, and 4.1 Å.

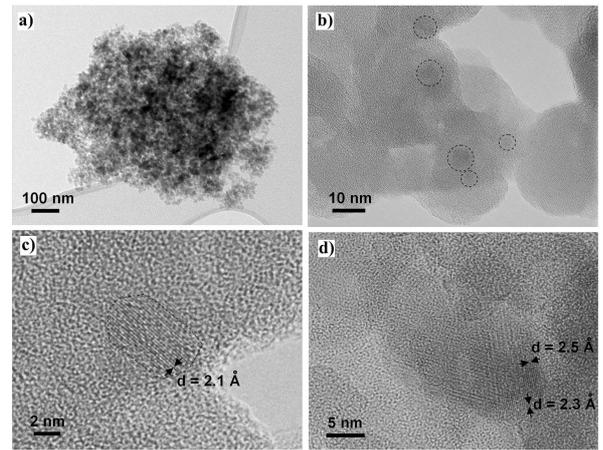


Fig. 2. TEM and HRTEM images of the sample with 3 at.% Mn: a) agglomerate of the amorphous silica particles, b) individual nanocrystalline particles dispersed in the amorphous silica matrix are marked with circles, c) enlarged area showing the nanocrystalline particle marked with a circle in the upper part of Fig. 2b), d) HRTEM images of the larger crystalline area composed of several smaller Mn-rich nanoparticles inside the amorphous silica matrix.

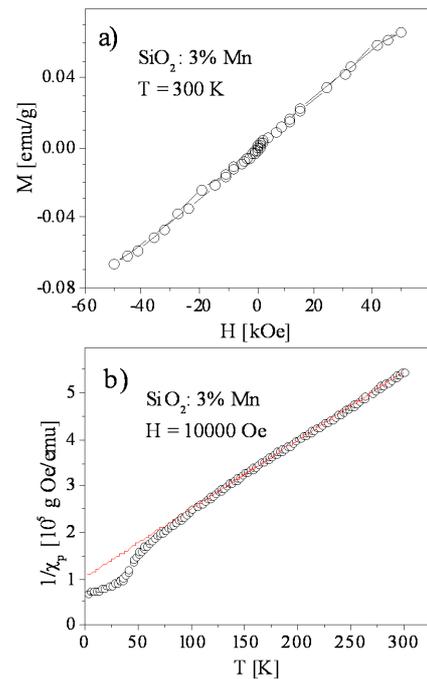


Fig. 3. a) Magnetic field dependent magnetization for the sample with 3 at.% Mn at $T = 300$ K, b) inverse high-temperature paramagnetic susceptibility for the sample with 3 at. % Mn at the magnetic field of 10000 Oe. The solid curve represents fit of the dependence (1) to the experimental data.

High-temperature magnetization measured as a function of magnetic field at different temperatures in the range 100–300 K for the sample with 3 at.% Mn shows that this dependence is not strictly linear. The M versus H dependence at high T exhibits hysteresis behavior, as can be seen in Fig. 3a which presents the $M(H)$ curve at $T = 300$ K. We have attributed this hysteresis behavior to a ferromagnetic (FM) component, so that the total magnetization can be described as $M(T, H) = \chi_p(T)H + M_{FM}$, where $\chi_p(T)$ is the paramagnetic susceptibility, H is the magnetic field and M_{FM} is the high-temperature ferromagnetic component of magnetization. The ferromagnetic component M_{FM} was estimated as intercept of the linear part of the total magnetization M . Paramagnetic susceptibility was then obtained by subtraction of ferromagnetic component from the total magnetization. The paramagnetic susceptibility of the sample with 3 at.% Mn so evaluated and corrected for diamagnetic susceptibility of silica is shown in Fig. 3b for the field of 10000 Oe. We find that $\chi_p(T)$ follows the Curie-Weiss dependence,

$$\chi_p(T) = \frac{C}{T - \theta}, \quad (1)$$

where C is the Curie-Weiss constant and θ is the Curie-Weiss temperature. The best fit parameters for the sample with 3 at.% Mn in the temperature range 200–300 K are obtained as $C = 6.8 \times 10^{-4}$ emu K/g_{sample}Oe which corresponds to $C = 1.36$ emu K/mol(Mn)Oe and $\theta = -(70 \pm 5)$ K indicating AFM interactions between the Mn ions. The FM component of magnetization observed in the high-temperature region is of the order of 1×10^{-3} emu/g_{sample} and is estimated to be about 5% of the total magnetization at the field of 10000 Oe. Until now room temperature FM has been observed in nonmagnetic oxide nanoparticles such as CeO₂, Al₂O₃, ZnO [9]. It was assumed that the origin of the room temperature FM in the oxide nanoparticles might be the exchange interaction between localized electron spins resulting from oxygen vacancies at the surface of the nanoparticles. Thin films of HfO₂ have been also reported to exhibit the room temperature FM in the absence of any doping. However, there is a conflicting report attributing the ferromagnetism in HfO₂ to possible iron contamination [10]. Origin of the high-temperature FM component of the magnetization in our sample SiO₂: 3 at.% Mn is unclear. We cannot exclude the possibility that some sort of FM impurity was present in the starting material used for the silica sol synthesis. The effective magnetic moment per Mn ion estimated from the Curie constant $C = N\mu_{\text{eff}}^2/3k_B$ is found to be $\mu_{\text{eff}} = 3.3\mu_B/\text{Mn}$. The values of C and μ_{eff} obtained for the sample with 3 at.% Mn are somewhat smaller than the theoretical spin-only values for Mn⁴⁺ with spin $S = 3/2$, $C(\text{Mn}^{4+}) = 1.87$ emu K/mol(Mn)Oe and $\mu_{\text{eff}}(\text{Mn}^{4+}) = 3.87\mu_B$. Calculated spin-only values of C and μ_{eff} for the Mn³⁺ and Mn²⁺ ions are considerably higher than those obtained in our experiment. The experimental values of C and μ_{eff} obtained for the 3

at.% Mn sample suggest that dominant magnetic phase at high-temperatures has a form of manganese dioxide. Tetragonal β phase of MnO₂ exhibits a phase transition at $T_N \approx 92$ K with AFM ordering below T_N , and Curie-Weiss temperature dependence of the magnetic susceptibility above 200 K with the Curie-Weiss temperature of -792 K [11]. Tetragonal α phase and orthorhombic γ phase of MnO₂ obey the classical Curie-Weiss law above 150 K with the Curie-Weiss temperature of about -300 K and -180 K respectively [12, 13]

Magnetic characteristics of the sample with 3 at.% Mn observed at high-temperatures resemble the properties of λ -MnO₂. There is no magnetic phase transition above 50 K, as can be seen in Fig. 4, and the Curie-Weiss temperature $\theta = -70$ K is much lower than that detected in the α , β and γ phase of MnO₂. The XRD spectrum of the 3 at.% Mn sample does not exhibit the XRD lines corresponding to λ -MnO₂ (Fig. 1). The HRTEM images, however, enabled us to measure several interplanar lattice spacings. Taking into consideration the limited accuracy of measurements, the spacings can be attributed to the cubic, spinel-type crystal structure of λ -MnO₂: 2.47 Å matching the d -value of (311) planes, 2.3 Å matching the d -value of (222) planes, 2.1 Å matching the d -value of (400) planes and 4.1 Å matching the d -value of (200) planes.

λ -MnO₂ form of manganese dioxide has been studied so far by several authors. The λ -MnO₂ bulk sample with composition Li_{0.10}MnO₂ prepared by acid treatment of LiMn₂O₄ showed the Curie-Weiss behavior in the high-temperature susceptibility with the Curie constant 1.97 emu K/mol(Mn)Oe ($\mu_{\text{eff}} = 3.97\mu_B$) and Curie-Weiss temperature of 104 K [5]. In the λ -MnO₂ sample with composition Li_{0.07}Mn₂O₄ prepared by electrochemical cycling of LiMn₂O₄ the high-temperature susceptibility exhibited the Curie-Weiss dependence with the Curie-Weiss temperature of -74 K and Curie constant 3.23 emu K/mol(Mn)Oe ($\mu_{\text{eff}} = 3.60\mu_B$) [6], which is also below the theoretical spin-only value of $3.87\mu_B$ for Mn⁴⁺. For our sample with 3 at.% Mn, where the presence of Li was not detected, the possible reason of the lower value of the parameter C and consequently of μ_{eff} compared to the spin-only value for Mn⁴⁺ could be somewhat higher value of manganese concentration determined by atomic absorption method (3 at.% Mn) compared to that actually present in this sample.

In contrast to the results of Greedan et al. [5], who observed long range AFM order by neutron diffraction below $T_N = 32$ K in their λ -MnO₂ sample which was also detected as a maximum in the magnetic susceptibility around 30 K, in our samples with 0.7 and 3 at.% Mn there is no peak in the ZFC and FC magnetization in the vicinity of 30 K (Fig. 4). Better evidence for the absence of the critical temperature near 30 K can be seen in the first derivative of the product MT with respect to temperature, $\partial(MT)/\partial T$, where M is the ZFC magnetization, which is presented in Figs. 5 and 6 for the samples with 0.7 and 3 at.% Mn respectively. There is no peak

in the $\partial(MT)/\partial T$ versus T plot in the vicinity of 30 K for both the samples. However, we observe another specific property in the samples at low-temperatures which can be seen in a form of weak peaks at about 9 K in the ZFC magnetization of the samples under the field of 1000 Oe (Fig. 4), as a peak at 8.4 K in the $\partial(MT)/\partial T$ versus T plot for the 0.7 at.% Mn sample (Fig. 5), and as a very weak peak at 8.7 K in the $\partial(MT)/\partial T$ versus T dependence for the 3 at.% Mn sample (Fig. 6). The low-temperature maximum appears to be magnetic field dependent which can be seen from the position of the peaks in the ZFC magnetization for the sample with 3 at.% Mn detected at 9 and 5.7 K under the fields of 1000 and 10000 Oe respectively (Fig. 4, inset).

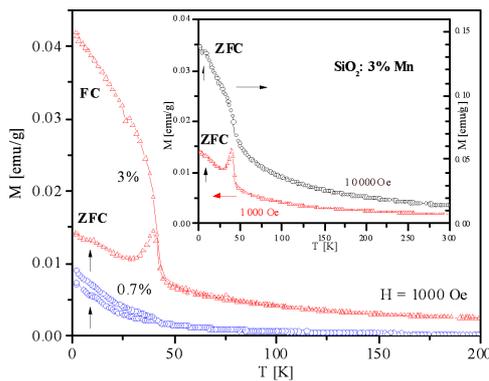


Fig. 4. Temperature dependence of the magnetization under the ZFC and FC conditions for the 0.7 and 3 at.% Mn samples at the magnetic field of 1000 Oe. The arrows indicate the positions of the low-temperature peaks in the ZFC magnetization at about 9 K for both the samples. Inset: ZFC magnetization as a function of temperature for the sample with 3 at.% Mn measured at the magnetic fields of 1000 and 10000 Oe: The arrows in the inset indicate the positions of the low-temperature peaks recorded at 9 and 5.7 K under the fields of 1000 and 10000 Oe respectively.

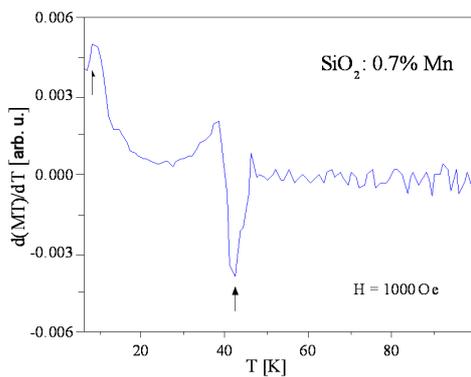


Fig. 5. $\partial(MT)/\partial T$ for the sample with 0.7 at.% Mn at the magnetic field of 1000 Oe. The position of the minimum at 42.4 K and of the low-temperature peak at 8.4 K are denoted by arrows.

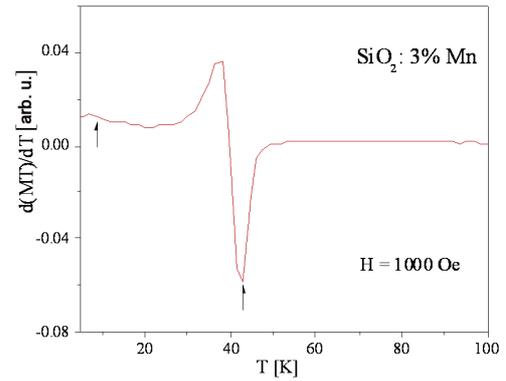


Fig. 6. $\partial(MT)/\partial T$ for the sample with 3 at.% Mn at the magnetic field of 1000 Oe. The position of the minimum at 43 K and of the low-temperature peak at 8.7 K are denoted by arrows.

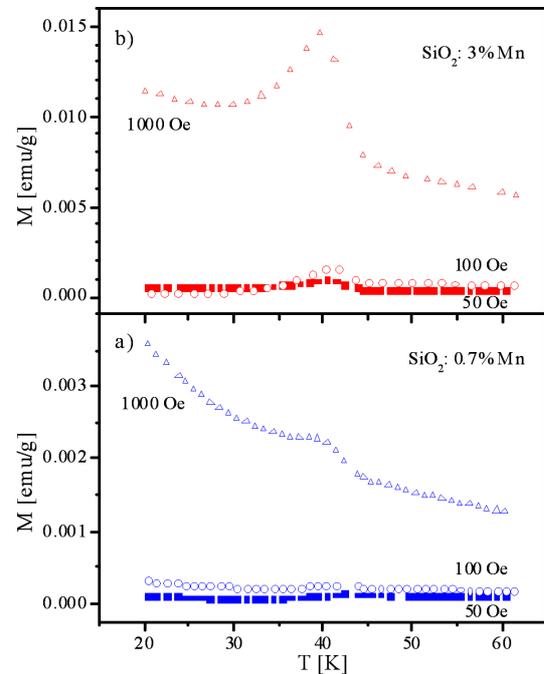


Fig. 7. ZFC magnetization versus temperature at the magnetic fields of 50, 100 and 1000 Oe for the samples with: a) 0.7 at.% Mn, b) 3 at.% Mn.

The real part χ' of the AC magnetic susceptibility for the 3 at.% Mn sample measured as a function of temperature at different frequencies exhibits a broad maximum of low intensity centered at about 14 K (Fig. 8). Appearance of the peaks in the ZFC magnetization and in the $\partial(MT)/\partial T$ versus T dependence between 5 and 10 K for both the samples, as well as a broad maximum in the real part of the AC susceptibility observed for the 3 at.% Mn sample at about 14 K could be attributed to a spin glass freezing of the Mn magnetic moments in the λ -MnO₂ phase of the samples. These properties are

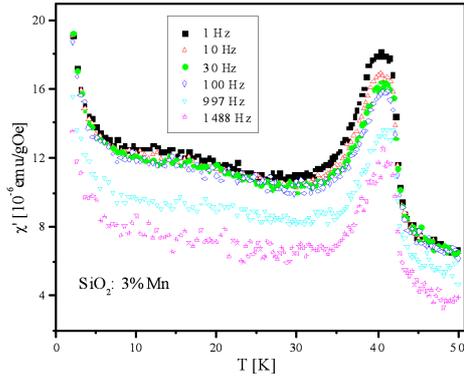


Fig. 8. Temperature dependence of the real part χ' of the AC magnetic susceptibility for the sample with 3 at.% Mn at different frequencies in the range 1-1488 Hz.

consistent with geometrical frustration of the manganese sublattice in the λ - MnO_2 crystal structure [5]. Spin glass behavior was also observed in the λ - $\text{Li}_{0.07}\text{Mn}_2\text{O}_4$ sample at temperatures below 16 K [6].

The ZFC and FC magnetizations of the samples measured at 1000 Oe, presented as a function of temperature in Fig. 4 clearly bifurcate below 43 K. At lower magnetic fields the ZFC curves exhibit a peak structure with a maximum dependent on the magnetic field. At higher magnetic field, $H = 10000$ Oe, the ZFC magnetization loses its peak structure and increases continuously with decrease in temperature (Fig. 4, inset). The ZFC magnetization versus temperature for the 0.7 at.% Mn sample detected at 50, 100 and 1000 Oe exhibits the maxima at 42.4, 41.4 and 38.4 K respectively (Fig. 7a). Fig. 7b presents the ZFC magnetization versus temperature for the 3 at.% Mn sample recorded at 50, 100 and 1000 Oe showing the maxima at 40.4, 40.3 and 39.7 K respectively. As can be seen, there is a systematic decrease of the temperature of the maximum in the ZFC magnetization for both the samples in the vicinity of 40 K with increase in magnetic field. The first derivative $\partial(MT)/\partial T$ for the 3 at.% Mn sample measured at 1000 Oe (Fig. 6) exhibits a maximum and a minimum at the positions (38.4 ± 0.5) K and (43.0 ± 0.5) K respectively and these characteristic temperatures were recorded within the limit of accuracy in measurements for all applied magnetic fields. The position of the minimum at 43 K corresponds to the Curie temperature of the bulk tetragonal Mn_3O_4 , $T_{C(\text{bulk})} = 42$ K [14], as well as of Mn_3O_4 nanoparticles of different sizes where the Curie temperature was found to vary from 40.5 K up to 44.3 K [15–17]. The first derivative $\partial(MT)/\partial T$ for the 0.7 at.% Mn sample measured at 1000 Oe (Fig. 5) shows a maximum and a minimum at the positions (38.4 ± 0.5) K and (42.4 ± 0.5) K respectively.

The temperature dependence of the real component of the AC magnetic susceptibility for the sample with 3 at.% Mn measured in the frequency range 1-1488 Hz is shown in Fig. 8. The real component χ' has a maxi-

um at a temperature T_m which has no noticeable shift with increasing frequency for lower frequencies up to 30 Hz where $T_m = 40.4$ K. For higher frequencies up to 1488 Hz T_m is shifted towards higher temperatures for about 0.3 K. Magnetic field dependent temperature of the maximum in the ZFC magnetization for both the samples and frequency dependent real component of the AC magnetic susceptibility of the 3 at.% Mn sample, which contains smaller Mn-rich nanoparticles with sizes 3-10 nm, suggest that the temperature T_m could be the blocking temperature of the Mn_3O_4 nanoparticles which appears just below the Curie temperature $T_C = 43$ K. Thermally activated blocking process of the magnetic moments of the small Mn_3O_4 nanoparticles in our samples is quite possible because superparamagnetic threshold volume at $T = 43$ K is $V_{th} = 1.5 \times 10^{-19}$ cm³, and for spherical particles this volume corresponds to the particle diameter $D_{th} \approx 7$ nm, as was already reported in Ref. 16. In distinction from our result, in the recently published paper [18] the behavior of an ensemble of the Mn_3O_4 nanoparticles is interpreted as a superparamagnetic blocking with the blocking temperature which can be, depending on the particle size, considerably lower than T_C .

A comparison of the FM component of magnetization in our sample with 3 at.% Mn obtained from the $M(H)$ data at 30 K (not shown), where it was possible to subtract the linear component from the total magnetization, with the literature data [14, 15] indicated the content of Mn_3O_4 in this sample to be about 0.2 wt.%. Taking into account the weight percent of manganese in the 3 at.% Mn sample determined by atomic absorption method as 2.734 wt.%, we estimated the content of the λ - MnO_2 phase to be about 4 wt.%. It is obvious that ferrimagnetism of the small amount of Mn_3O_4 is strong enough at lower temperatures to mask the relatively weak magnetism of λ - MnO_2 with AFM interactions. On the other hand, paramagnetic component of the magnetization originating from λ - MnO_2 overcomes paramagnetic contribution of the Mn_3O_4 phase observed in the high-temperature magnetic susceptibility.

4. Conclusion

Magnetic properties of the samples with 0.7 and 3 at.% Mn incorporated in an amorphous SiO_2 matrix have been studied. TEM analysis has shown that the samples contain agglomerates of amorphous silica particles 10-20 nm in size. In the amorphous matrix two types of nanocrystalline Mn-rich particles are dispersed: smaller Mn-rich nanocrystals with dimensions between 3 and 10 nm, and larger crystalline areas consisting of aggregates of the smaller nanoparticles. At higher temperatures dominant magnetic phase is λ - MnO_2 . At temperatures below $T_C = 43$ K strong ferrimagnetism originating from the minor Mn_3O_4 phase masks the relatively weak magnetism of λ - MnO_2 with AFM interactions. Magnetic field dependent ZFC magnetization of both the samples and

frequency shift of the real component of the AC susceptibility in the 3 at.% Mn sample suggest that the magnetic moments of the smaller Mn₃O₄ nanoparticles are subject to thermally activated blocking process just below the Curie temperature T_C . Appearance of the maximum in the ZFC magnetization of the samples below 10 K indicates spin glass freezing of the magnetic moments at low temperatures which probably occurs in the geometrically frustrated Mn sublattice of the λ -MnO₂ crystal structure.

Acknowledgments

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References

- [1] G.H. Lee, S.H. Huh, J.W. Jeong, B.J. Choi, S.H. Kim, H.C. Ri, *J. Am. Chem. Soc.* **124**, 12094 (2002) and references therein.
- [2] M.M. Thackeray, *Prog. Solid St. Chem.* **25**, 1 (1997).
- [3] M. Wakihara, *Materials Science and Engineering* **R33**, 109 (2001).
- [4] P.G. Bruce, B. Scrosati, J.M. Tarascon, *Angew. Chem. Int. Ed.* **47**, 2930 (2008).
- [5] J.E. Greedan, N.P. Raju, A.S. Wills, C. Morin, S.M. Shaw, J.N. Reimers, *Chem. Mater.* **10**, 3058 (1998).
- [6] Young-II Jang, B. Huang, F.C. Chou, D.R. Sadoway, Y.M. Chiang, *J. Appl. Phys.* **87**, 7382 (2000).
- [7] A.S. Wills, N.P. Raju, J.E. Greedan, *Chem. Mater.* **11**, 1510 (1999).
- [8] J.R. Martínez, S. Palomares-Sánchez, G. Ortega-Zarzosa, F. Ruiz, Y. Chumakov, *Materials Letters* **60**, 3526 (2006).
- [9] A. Sundaresan, R. Bhargavi, N. Rangarajan, U. Sidesh, C.N.R. Rao, *Phys. Rev. B* **74**, 161306(R) (2006).
- [10] D.W. Abraham, M.M. Frank, S. Guha, *Appl. Phys. Lett.* **87**, 252502 (2005).
- [11] F. Luo, W. Song, C.H. Yan, *Chem. Phys. Lett.* **431**, 337 (2006).
- [12] A.M.A. Hashem, H.A. Mohamed, A. Bahloul, A.E. Eid, C.M. Julien, *Ionics* **14**, 7 (2008).
- [13] N. Wang, Y. Gao, J. Gong, X. Ma, X. Zhang, Y. Guo, L. Qu, *Eur. J. Inorg. Chem.* 3827 (2008).
- [14] K. Dwight, N. Menyuk, *Phys. Rev.* **119**, 1470 (1960).
- [15] N. Wang, L. Guo, L. He, X. Cao, C. Chen, R. Wang, S. Yang, *Small* **3**, 606 (2007).
- [16] E. Winkler, R.D. Zysler, D. Fiorani, *Phys. Rev. B* **70**, 174406 (2004).
- [17] Y.Q. Chang, X.Y. Xu, X.H. Luo, C.P. Chen, D.P. Yu, *J. Cryst. Growth* **264**, 232 (2004).
- [18] R.J. Tackett, J.G. Parsons, B.I. Machado, S.M. Gaytan, L.E. Murr, C.E. Botez, *Nanotechnology*, **21**, 365703 (2010).