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Electromagnetic Properties of Substituted Barium Hexaferrites

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Substituted M-type ferrite $BaFe_{12-2x}(Me_1Me_2)_xO_{19}$, $0 \le x \le 0.6$ was prepared by both mechanical alloying and precursor method, where $Me_1 =$ Co, Ni, Zn, Sn and Me₂ =Ru, Ti, Zr, Sn. Magnetic phase purity, change of saturation polarization, Curie temperature, coercivity and magnetic susceptibility was studied as function of x. Attention was focused to results obtained for $(NiRu)_x$, $(ZnRu)_x$, and $(SnRu)_x$ mixtures with low doping ratio x.

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1. Introduction

M-type Ba-hexaferrite exhibits high specific saturation polarization J_{s-m} , strong uniaxial crystalline anisotropy K_1 and large coercivity H_c . The strong uniaxial anisotropy of Ba-hexaferrite can be reduced by substitution of iron by other cations or mixture of cations. One aim of such substitution was to decline the magnetization of one of the sublattices from antiferromagnetic axis, i.e. to decline it from collinear and anticollinear ordering, respectively. The magnetocrystalline anisotropy becomes planar for a critical doping ratio x_c in some cases; these planar hexaferrites therefore become to be interesting materials for magnetic recording and microwave absorption devices. It is an efficient way for controlling the value of H_c , while preserving the value of J_{s-m} . Various substitutions are possible in BaFe₁₂O₁₉ composition — big divalent cations on Ba²⁺ site and small trivalent cations, or divalent Me₁²⁺ together with tetravalent Me₂⁴⁺ ions combinations on the

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 ${\rm Fe}^{3+}$ sites [1–5]. Substitutions of Fe ions have a big influence, directly affecting the magnetic structure.

The magnetic structure and properties of substituted M-type Ba-hexaferrites with composition $BaFe_{12-2x}(Me_1Me_2)_xO_{19}$ were studied. Divalent $Me_1 = Co$, Ni, Sn, Zn and tetravalent $Me_2 = Ru$, Ti, Zr, Sn ion combinations were used in various compounds. The phase purity of the samples was evaluated by temperature dependence of initial susceptibility $\chi(T)$ and specific saturation polarization $J_{s-m}(T)$ [6] and it was verified by the Mössbauer spectroscopy and X-ray diffraction. The particle size and shape was observed under a transmission electron microscopy (TEM). The particles of all ferrite samples exhibit a plate-like, nearly hexagonal shape with diameter between 0.3 and 1.1 μ m.

2. Theory

The intrinsic properties arise from the M-crystal structure — five distinct Fe-sublattices, which are coupled by super-exchange allowing only parallel (\uparrow) or antiparallel (\downarrow) orientation. The magnetic moments of the iron ions are arranged in parallel to hexagonal *c*-axis, but with opposite spin directions of the sublattices. Their mutual orientation is given by the known Gorter model: 12k (\uparrow), 4f₁ (\downarrow), 4f₂ (\downarrow), 2a (\uparrow) and 2b (\uparrow). Three sites named 12k, 2a, and 4f₂ have octahedral coordination, 4f₁ has tetrahedral and 2b site has a fivefold (bipyramidal) coordination. The resulting theoretical magnetic polarization *J* at a temperature *T* of BaFe₁₂O₁₉ per formula unit (f.u.) can be approximated by simple summation according to formula

 $J(T) = 6m_{12k}(T) - 2m_{4f1}(T) - 2m_{4f2}(T) + 1m_{2a}(T) + 1m_{2b}(T),$ (1) where m_n is the magnetic moment of Fe³⁺ ion in *n*-th sublattice. Assuming a magnetic moment of $5\mu_{\rm B}$ (Bohr magneton) per Fe³⁺ ion at 0 K we get a net polarization of $20\mu_{\rm B}$ per f.u. for the pure BaFe₁₂O₁₉ at this temperature. It is in agreement with the observed saturation polarization for the pure compound $J_{\rm s}$. According to Eq. (1) one can expect a change of the net magnetic polarization when the Fe³⁺ ions in the spin-down sublattices (4f₁,4f₂) and/or spin up sublattices (12k, 2a, 2b) are substituted.

3. Results

Influence of substitution to net polarization, magnetostructural anisotropy and coercivity H_c was investigated as well as possible limits of such approach. Change of magnetic polarization with doping ratio x is expressed by the difference $\Delta J_{s-m} = J_{s-m}(x) - J_{s-m}(0)$, where $J_{s-m}(0)$ is the specific magnetic polarization of pure BaM ferrite (x = 0). This dependence for selected samples prepared by mechanical alloying (Mx) is shown in Fig. 1a. It can be seen that ΔJ_{s-m} increases with x for ZnRu, NiRu, and SnRu substitutions presumably, because the magnetic moments in the spin-down sublattices (4f₁,4f₂) decrease more than those in the spin-up sublattices (2a,2b) (Eq. (1)). The Ru⁴⁺ ions seem to increase ΔJ_{s-m} ,

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Fig. 1. Change of specific polarization ΔJ_{s-m} (a) and coercivity ΔH_c (b) as a function of x for samples (Mx).

it can also be attributed to a strong preference of Zn^{2+} to tetrahedral, Ni²⁺ and Sn^{2+} to octahedral coordination and no particular preference Ru^{4+} ions in Me_1Ru mixtures.

The dependence of coercivity on x is expressed by coercivity difference $\Delta H_{\rm c}(x) = H_{\rm c}(x) - H_{\rm c}(0)$, where $H_{\rm c}(0)$ is a coercivity of sample with x = 0. Dependence $\Delta H_{\rm c}(x)$ for selected $({\rm Me}_1, {\rm Me}_2)_x$ substituent ions is shown in Fig. 1b. The strong fall of $\Delta H_{\rm c}$ near x = 0.1 and slight decrease in $\Delta H_{\rm c}$ observed near x = 0.3 for $({\rm NiRu})_x$ and $({\rm ZnRu})_x$ leads to the conclusion that the reached value of $H_{\rm c}$ starts to correspond with planar anisotropy. This value of $H_{\rm c}$ change $(x \approx 0.3)$ is much lower than those of $x \approx 0.3 \div 0.6$ necessary for the $H_{\rm c}$ change of other $({\rm Me}_1, {\rm Me}_2)_x$ ions. In order to understand the decrease in $\Delta H_{\rm c}$ with increasing x, ionic distribution in sites should be taken into account. The fivefold site 2b has the strongest influence on the high anisotropy of BaFe₁₂O₁₉ ferrite. Structural refinements on $({\rm NiRu})_x$ and $({\rm ZnRu})_x$ substitutions show that ${\rm Ru}^{4+}$ preferentially substitutes Fe³⁺ ions on 2b site for low x, and then on octahedral sites $4f_2$ and 2a (for Ni) or on the tetrahedral sites $4f_1$ (for Zn).

Therefore, we attribute the rapid change of anisotropy and H_c in $(NiRu)_x$ and $(ZnRu)_x$ compounds to preferential occupation of 2b site by Ru^{4+} cations. Moreover, we suppose that the nature of divalent cation — Ni²⁺ (3d⁸) or Zn²⁺ (3d¹⁰) has influence but lower on the anisotropy change. This opinion is confirmed by the fact that the cationic distribution is the same in $(NiRu)_x$ as in $(Sn^{2+}Ru)_x$, but ΔH_c values of $(SnRu)_x$ are significantly lower as compared with $(NiRu)_x$ and $(ZnRu)_x$. It is probably due to Ni²⁺ or Zn²⁺ having 3d⁸ or 3d¹⁰ orbitals and Sn^{2+} having 3d¹⁰, 4d¹⁰ orbitals. This opinion could be supported by the fact, that $(CoRu)_x$ substitution [4] $(Co^{2+} has 3d^7 orbitals)$ has the same effect on anisotropy as $(NiRu)_x$. In addition, T_c decreases with the substitution rate xfrom 442°C (x = 0) down to 426°C for $(ZnRu)_{0.3}$, 435°C for $(NiRu)_{0.3}$ and 420°C for $(CoRu)_{0.3}$. Significantly lower value of $T_c = 389°C$ was found in the case of $(SnRu)_{0.3}$ substitution.

Furthermore presence of secondary phases containing spinel ferrite was found in $\chi(T)$ curves at temperature above T_c of the hexagonal phase for NiMe₂ subJ. Sláma et al.



Fig. 2. Temperature dependence of magnetic susceptibility $\chi(T)$ of NiMe₂⁴⁺ prepared by Mx (a) and Sk (b).

stituted samples Me₂ = Ti, Zr, Sn, Ru) prepared only by mechanical alloying (Fe/Ba = 10.0), Fig. 2a. This phase was not detected by X-ray analysis; it is most likely the NiFe₂O₄ (or magnetite Fe₃O₄) compound, with $T_c \approx 585^{\circ}$ C. On the other hand, no traces of spurious phases were found in BaM samples synthesized by the citrate precursor method (Sk) (Fe/Ba = 10.8) with x up to 0.4, Fig. 2b. Ni spinel ferrite was obtained by Turilli et al. [1] in Sr(NiTi) M-hexaferrite prepared by the metalorganic precursor method. Small doping rate of x was sufficient in some cases - to achieve nearly planar anisotropy in (NiRu)_{0.3} and (ZnRu)_{0.3} mixtures with a low temperature coefficient of H_c . Such ferrites could be very promising media for microwave devices and absorption at higher temperatures.

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